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SECTION 12.1 THE ARITHMETIC OF EQUATIONS

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Guided Practice: Stoichiometry Mass to Mass Problems To convert from mass in grams of a reactant to volume, in liters, of a product (reverse the process for liters to grams): • Use factor label method • Use mass of reactant from the Periodic Table 1 mol=_____ g • Use the mole to mole ratio from the balanced reaction

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Stoichiometry 353 12.1 FOCUS Objectives 12.1.1 Explain how balanced equa-tions apply to both chemistry and everyday situations. 12.1.2 Interpret balanced chemical equations in terms of moles, representative particles, mass, and gas volume at STP. 12.1.3 Identify the quantities that are always conserved in chemical reactions. Guide for Reading

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12.1 The Arithmetic of Equations 12

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