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H 2 O 18.02 g/mol CH 4 16.04 g/mol H 2 CO 3 46.03 g/mol C 6 H 12 O 6 180.15 g/mol
Activity B: Canceling units Get the Gizmo ready: Check that the equation is still: $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
Introduction: While solving problems in stoichiometry, it is useful to pay attention to the units of the answer.

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Chapter 12 Assessment Stoichiometry Answer Key

Chapter 12 Review Stoichiometry Important Vocabulary • Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles

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Chapter 12 Stoichiometry Section Review Chapter 12 REVIEW: Stoichiometry, Theoretical, Actual & Percent yield Part I. Stoichiometry 1. 1 ZnI₂ + 1 Zn + 1 I₂
How many grams of iodine will you produce if you begin your rxn with 56.7g of zinc iodide? 56.7 g 1 mole 1 mole 253.8 g 319.19 g 1 mole 1 mole = 45.08 grams of Iodine 2. 1 Pb₂(CO₃)₃ + 6 K ...

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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ a. If you use 2.3 moles of Al, how many moles of Al₂O₃ can you make? b. If you want 3.9 moles of Al₂O₃, how many moles of O₂ are needed? 2.

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Stoichiometry, by definition, is the calculation of the quantities of reactants or products in a chemical reaction using the relationships found in the balanced chemical equation. The word stoichiometry is actually Greek from two words: $\sigma\tau o\iota\chi\iota\mu\epsilon\tau\alpha$ (\sigmaigma \tauau \omicron micron \iotaota \kappaappa \etata \iotaota \omicron micron \nuu \), which means "element", and $\mu\epsilon\tau\rho$ (\muu \epsilonpsilon lon \tauau \rho ho \omicron micron \nuu), which means "measure".

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