

## 8 Covalent Bonding Answers Core Teaching Resources

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In the Covalent Bonding game, players bond together atoms to create target molecules with unique bond polarities and molecular shapes. Check out our Free Teacher Resources here! Take a peek inside the Covalent Bonding game for a brief overview of the concepts covered through gameplay.

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1) Ionic bonds-transfer of electrons. (gained or lost; makes formula unit) 2) Covalent bonds-sharing of electrons. The resulting particle is called a "molecule". Covalent Bonds. • The word covalent is a combination of the prefix co- (from Latin com, meaning "with" or "together"), and the verb. valere, meaning "to be strong".

Chapter 8 " Covalent Bonding "

242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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atom equally attracts the pair of shared electrons.

## 8.E: Chemical Bonding Basics (Exercises) - Chemistry ...

Covalent bonds are formed when electrons are shared between atoms. Only non-metal elements participate in covalent bonding. As in ionic bonding, each atom gains a full outer shell of electrons.

Covalent bonds between atoms are strong and the bonded unit of atoms is what is called a molecule.

Covalent Bonding Answers Worksheets - Learny Kids

~~How to Draw Covalent Bonding Molecules~~  
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Section Review Objectives • State a rule that usually tells how many electrons are shared to form a covalent bond • Describe how electron dot formulas are used • Predict when two atoms are likely to be joined by a double or a triple covalent bond • Distinguish between a single covalent bond and other covalent bonds

## 8.2 The Nature of Covalent Bonding 8

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the attachment between atoms is called  
covalent bond in a covelant bond, the  
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Based on relative electronegativities,  
 classify the bonding in each compound as  
 ionic, covalent, or polar covalent. Indicate  
 the direction of the bond dipole for each  
 polar covalent bond. NaBr; OF<sub>2</sub>; BCl<sub>3</sub>;  
 the S – S bond in CH<sub>3</sub>CH<sub>2</sub>SSCH<sub>2</sub>CH<sub>3</sub>;  
 the C – Cl bond in CH<sub>2</sub>Cl<sub>2</sub>; the O – H bond  
 in CH<sub>3</sub>OH; PtCl<sub>4</sub> 2 –

Section 8.2 The Nature of Covalent  
 Bonding 219 You can draw the electron  
 dot structure for ammonia (NH<sub>3</sub>), a

suffocating gas, in a similar way. The  
 ammonia molecule has one unshared pair  
 of electrons. Hydrogen atoms 3H. or  
 Nitrogen atom Ammonia molecule N.... .  
 H H.H.... H H H.N HH 1s 1s 1s 2s 2p N H  
 1s Ammonia molecule Methane (CH<sub>4</sub>)  
 contains four single covalent bonds. The  
 carbon atom