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$\ln K = (\Delta S^\circ/R) - (\Delta H^\circ/RT)$ The equilibrium for the dissociation of one mole of iodine vapour is 3.94×10^{-4} at 900K and the standard enthalpy is 154 kJmol⁻¹. Calculate the equilibrium constant at 1200K assuming the standard enthalpy remains constant. $\ln (K_2/K_1) = - \Delta H^\circ/R (1/T_2 - 1/T_1)$, ANS = 6.7×10^{-2} at 1200K.

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