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 $\times 10^{-9} \text{ M}^{-1} \text{ s}^{-1}$ a. \ln
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where k is the rate constant and c
is the initial concentration. This is
the equation that describes first-
order kinetics. If $1/\text{concentration}$
vs. time is linear, the data will fit
to the following equation: $1/[c] = kt + 1/[c]_0$. where k is the rate
constant and c is the initial
concentration.

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Unit 1 (Kinetics) - Chemistry 12

Write the rate expression in
terms of $[\text{reactant}] / \Delta t$ and $[\text{product}] / \Delta t$ for the reaction: (1)
 $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$.

Determine. The order of each
reactant, The order of the
reaction, Write the rate law and.

*Practice Kinetic And Potential
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Reaction Rates Page 2 b) If the rate of consumption of magnesium is 5.0×10^{-9} mol/s, find the rate of consumption of HCl in moles/s.

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Work in groups on these problems. You should try to

answer the questions without referring to your textbook. If you get stuck, try asking another group for help. 1. For the reaction below, if substance A is disappearing at a rate of $(1.82 \times 10^{-2} \text{ mol L}^{-1} \text{ s}^{-1})$, at what rate is C appearing? $[3 \text{ A} + 3 \text{ B} \rightarrow 5 \text{ C} + 2 \text{ D}]$ 2.

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students studying kinetic energy (part of the energy topic). This worksheet has been created with all abilities in mind, and has extension questions to stretch and challenge the more able. The answers have been provided. This resource hasn't been reviewed.

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Worksheet 2 1) A Reaction

Is First Order ...

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