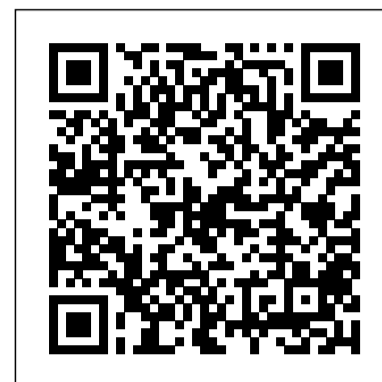


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Write the rate expression in terms of  $[reactant]/t$  and  $[product]/t$  for the reaction: (1)  $N_2 + 3H_2 \rightarrow 2NH_3$ . 2. Determine.

The order of each reactant, The order of the reaction, Write the rate law and.

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Work in groups on these problems. You should try to answer the questions without referring to your textbook. If you get stuck,

try asking another group for help. 1. For the reaction below, if substance A is disappearing at a rate of  $(1.82 \times 10^{-2}) \text{ mol L}^{-1} \text{ s}^{-1}$ , at what rate is C appearing?  $[3A + 3B \rightarrow 5C + 2D]$  2.

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where k is the rate constant and c is the initial concentration. This is the equation that describes first-order kinetics. If 1/concentration vs. time is linear, the data will fit to the following equation:  $1/[c] = kt + 1/[c]_0$ . where k is the rate constant and c is the initial concentration.

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[1-Reaction Kinetics Worksheet 1-1 Measuring Reaction Rates Page 2 b\) If the rate of consumption of magnesium is  \$5.0 \times 10^{-9} \text{ mol/s}\$ , find the rate of consumption of HCl in moles/s.](#)