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Consider the following reaction:  $\text{CuS(s)} + \text{O}_2\text{(g)} \rightleftharpoons \text{Cu(s)} + \text{SO}_2\text{(g)}$  A reaction mixture initially contains 2.9 M  $\text{O}_2$ . Determine the equilibrium concentration of  $\text{O}_2$  if  $K_c$  for the reaction at this temperature is 1.5.

[A reaction mixture initially contains 2.9 ... - Yahoo Answers](#)

Source(s): This was a real head scratcher, until I read in my chemistry book "Chemistry a Molecular Approach" by Nivaldo J. Tro, on page 257, the following: To get  $\Delta H^\circ_{\text{rxn}}$  per mole of a particular reactant - a more general quantity - you simply divide by the number of moles that actually reacted, and this is what gave me the idea that maybe I should be dividing the obtained moles by 4 ...

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