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ANSWERS. MODULE 1:

FUNDAMENTALS IN
CHEMISTRY 1. (a)(i) Define
the term bond energy [1. mark]

The bond energy represents:

The quantity of energy evolved
when the bond is formed. Also,
the bond energy is the amount
of energy absorbed when the
bond is broken. (ii) The bond
lengths and bond energies of
carbon-carbon single, double
and triple bonds are listed
below.

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there is a change in structure
from giant molecular to giant
metallic and a change in
bonding from covalent to
metallic. From C to Ge
elements exhibit a giant
molecular structure. Sn and
Pb exhibit a giant metallic
lattice structure.

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Explanations. 1. CAPE
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QUESTION 1 SOLUTION
AND EXPLANATIONS (a)

A dative covalent or
coordinate covalent bond
is one in which one of the
atoms supplies both the
shared electrons to the
covalent bond. A dative
covalent bond is written
A→B, where the direction
of the arrow shows the
direction in which the
electron pair is donated.

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