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SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287 – 296)

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SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287 – 296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287 – 289)  
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Chemistry 101 – Chemical Quantities  
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 • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the  
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10, that can be made from  $1.09 \times 10^4$  kilograms of phosphorus in the second of these three reactions. The answer is  $2.50 \times 10^4$  kg P 40 10. We used the following steps: Balance chemical equations. (Section 4.1.) Write or identify the definitions of

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solution, solute, and solvent. (Section 4.2) Given a description of a solution, identify the