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SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

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Chapter 10 "Chemical Quantities" Vocab. the SI unit representing 6.02 X 10^23 representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

Chemistry 101 - Chemical Quantities (Empirical/Molecular Formula)

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Chapter 10 Chemical Quantities Answers

Chapter 10 Chemical Quantities91 SECTION

10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.

the SI unit representing 6.02 X 10^23 representative particles of a substance: Avogadro's number: 6.02 X 10^23 particles: standard temperature and pressure (00 C, 1 atm) the temperature and pressure at which one mole of gas occupies a volume of 22.4 L: molar volume: volume of a gas that contains one mole of the gas, is 22.4 L at STP: Avogadro ...