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# Chapter 10 Chemical Quantities Test

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Section Review

Objectives •

Relate Avogadro's number to a mole of a substance •

Calculate the mass of a mole of any substance •

Describe methods of measuring the

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amount of something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles Chapter 10 Chemical Quantities (The Mole!!) Flashcards by ... Unit 4 discusses measurement in science as well as measurements that are specific to Chemistry. Chapter 3 deals with general math topics including significant digits

and scientific notation. Chapter 10 introduces the idea of the mole and the use of equivalences to the mole as conversion factors. Objectives Vocabulary Key Equations Test and improve your knowledge of Prentice Hall Chemistry Chapter 10: Chemical Quantities with fun multiple choice exams you can take online with Study.com Chapter 10 Chemical Quantities Test CHAPTER 10:

Chemical Quantities. CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the Chemistry Chapter 10 Chemical Quantities Test Flashcards ... Start studying chapter 10 chemical quantities test. Learn vocabulary, terms, and more with flashcards, games, and other

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composition: a description of the relative amounts of each element in a compound:

empirical formula: the lowest whole-number ratio of the atoms of the elements in a ...

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The mole was founded by a scientist named Avagadro, and he decided to use the

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The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

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Honors Chemistry

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