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## Chapter 10 Chemical Quantities Test

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CHAPTER 10: Chemical Quantities. CHAPTER 10: Chemical Quantities BASICS:

- The basic unit that is used to determine the amount of a chemical substance is called a mole
- A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance
- The mole was founded by a scientist named Avagadro, and he decided to use the

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account; E-mail to a friend ... Use these activities to learn the vocabulary presented in this chapter. A B; percent composition: a description of the relative amounts of each element in a compound: empirical formula: the lowest whole- number ratio of the atoms of the elements in a ...

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Unit 4 discusses measurement in science as well as measurements that are specific to Chemistry.

Chapter 3 deals with general math topics including significant digits and scientific notation. Chapter 10 introduces the idea of the mole and the use of equivalences to the mole as conversion factors.

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*Objectives Vocabulary Key Equations*

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Chapter 11 - Chemical Reactions Chapter 12 - Stoichiometry Chapter 13 - States of Matter Chapter 14 - Behavior of Gases Chapter 15 - Water and Aqueous Systems ... Chapter 10 - Chemical Quantities. Test Review  
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Chapter 10: Chemical Quantities- Chapter Test B (pages 256-259) by Pearson Education Learn with flashcards, games, and more – for free.

Unit 4 - Chapters 3 & 10 - Mrs. Gingras' Chemistry Page

Chapter 10 (Chemical Quantities) Test Study Guide The mole is the SI unit used to measure the number of representative particles in a substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

Quia - Chapter 10 "Chemical Quantities"

Vocab

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**Chemistry I H: Chapter 10 Chemical Quantities- Chapter Test B**

Chapter 10 Chemical Quantities 241 Section Review Objectives • Relate Avogadro's number to a mole of a substance • Calculate the mass of a mole of any substance • Describe methods of measuring the amount of something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles

**Prentice Hall Chemistry Chapter 10: Chemical Quantities ...**

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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Honors Chemistry ... Chapter 10 - Chemical Quantities. 10.1 The Mole: A Measurement of Matter 10.2 Mole - Mass and Mole - Volume Relationships 10.3 Percent Composition and Chemical Formulas Chapter 10 Test Review 10.2 Practice Problems ANS. Chapter 11 - Chemical Reactions.

**Chapter 10 Chemical Quantities Answer Key**

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