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Chapter 10 – Chemical Quantities

CHEMICAL QUANTITIES SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287 – 296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287 – 289)

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Quantities (Empirical/Molecular Formula) ~~Moles, Atoms, and Molecules~~ Introduction to ~~Quantities Review~~ Introduction to Chemical  
 Step by Step Stoichiometry Practice Problems Limiting Reactant and Excess Reactant How Quantities (the mole and molar mass)  
 | How to Pass Chemistry Chemical Quantities to Balance Chemical Equations in 5 Easy Chemical Quantities - The Mole Unit 10 Part  
Review Introduction to Chemical Quantities Steps: Balancing Equations Tutorial Chemical 2 ~~CH 10 CHEMISTRY THE MOLE IN~~  
 (the mole and molar mass) Chemical Quantities and Reactions, part 5 - Grams to ~~REVIEW~~ Stoichiometry Basic Introduction,  
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Unit 7: Chemical Quantities

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Chapter 10 Chemical Quantities Practice Problems

Worksheet ...

Chemical Quantities THE MOLE AND

QUANTIFYING MATTER 10.1 The Mole: A

Measurement of Matter Essential Understanding The

mole represents a large number of very small

particles. Reading Strategy Frayer Model The Frayer

Model is a vocabulary development tool. The center of the

CHAPTER 2/CHAPTER 10/CHAPTER 11

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Section 10.1 – The Mole: A Measurement of Matter. You often measure the amount of something by count, by mass, or by volume. A mole (mol) of a substance is  $6.02 \times 10^{23}$  representative particles of that substance.  $6.02 \times 10^{23}$  is called Avogadro ' s number. 1 mole =  $6.02 \times 10^{23}$  representative particles