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molecular theory.
particles of matter
are always in
motion. 1.
expansion 2.
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compressability.
expansion. 1.
gases do not have
a definite shape or
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hypothesis. States that
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at the same
temperature and
pressure contain equal
numbers of particles.
Avogadro's number.
the number of
representative particles
in a mole $6.02 \times$

10^{23} ...
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Mole". Terms in this
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Number. "The
number
 6.0221367×10^{23} ,
which is the number
of representative
particles in a mole,
and can be rounded
to three significant
digits 6.02×10^{23} ".
Empirical Formula.
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Allotrope. One or two or more

different molecular forms of an element in the same physical state; oxygen (O₂) and ozone (O₃) are

allotropes of the element oxygen.

Amorphous Solid. A term used to describe a solid that lacks an ordered internal structure; denotes a random

arrangement of atoms.

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Kinetic molecular theory. a theory that explains that the behavior of physical systems depends on the combined actions of the molecules constituting the system. Ideal gas. an imaginary gas whose particles are infinitely small and do not interact with each other.

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