

Chapter 10 Chemistry Vocab

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Chemistry 1405 Chapter 10 1. Stoichiometry – calculate the amount of reactants required for a chemical reaction to be complete, and the amount of products made after the reaction comes to an end. 2. Use of a chemical equation to obtain conversion factors – a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles.

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Avogadro's Law- a statement that the volume of a gas maintained at constant pressure is directly proportional to the number of moles of the gas. 10. Ideal gas equation-an equation for the gases that embody Boyle's Law, Charles's Law, and Avogadro's Hypothesis in the form $PV=nRT$ 11. You've reached the end of your free preview.

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A B; radioactivity: the process in which an unstable atomic nucleus emits charged particles and energy: radioisotope: any atom containing an unstable nucleus

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the SI unit representing 6.02×10^{23} representative particles of a substance: Avogadro's number: 6.02×10^{23} particles: standard temperature and pressure (00 C, 1 atm) the temperature and pressure at which one mole of gas occupies a volume of 22.4 L: molar volume: volume of a gas that contains one mole of the gas, is 22.4 L at STP:

Avogadro ...

Quia - Chapter 10 "Chemical Quantities" Vocab Chemistry 1405 Chapter 10 1. Stoichiometry – The quantitative relationships among the substances involved in a chemical reaction, established by the equation for the reaction. 2. Use of a Chemical Equation to Obtain Conversion Factors – Conversion between moles of one species and moles of another species in a chemical change.

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