

Chapter 10 Liquids And Solids Answers

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Chapter 10: Liquids and Solids. STUDY. PLAY. condensed states. the solid and liquid states when atoms or molecules have a limited amount of movement. intermolecular forces. forces of attraction between molecules. dipole-dipole attraction.

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In this chapter, the nature of these interactions and their effects on various physical properties of liquid

and solid phases will be examined. 10.0: Prelude to Liquids and Solids In the liquid and solid states, these

interactions are of considerable strength and play an important role in determining a number of physical

properties that do depend on the chemical identity of the substance.

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[Chapter 10 Liquids and Solids Figure 10.1 Solid carbon dioxide \(" dry ice " , left\) sublimates vigorously when](#)

[placed in a liquid \(right\), cooling the liquid and generating a fog of condensed water vapor above the cylinder.](#)

[\(credit: modification of work by Paul Flowers\) Chapter Outline 10.1 Inter molecular Forces 10.2 Properties of](#)

[Liquids 10.3 Phase Transitions](#)

[Chapter 10 - Liquids and Solids](#)

10.1: Prelude to Solids and Liquids Liquids flow when a small force is placed on them, even if only very slowly.

Solids, however, may deform under a small force, but they return to their original shape when the force is

relaxed. This is how glass behaves: it goes back to its original shape (unless it breaks under the applied force).

[Chapter 10 Liquids And Solids](#)

[Chapter 10. Liquids and Solids. Introduction; 10.1 Intermolecular Forces; 10.2 Properties of Liquids;](#)

[10.3 Phase Transitions; 10.4 Phase Diagrams; 10.5 The Solid State of Matter; 10.6 Lattice Structures in](#)

[Crystalline Solids; Chapter 11. Solutions and Colloids. Introduction; 11.1 The Dissolution Process; 11.2](#)

[Electrolytes; 11.3 Solubility; 11.4 Colligative Properties](#)

[List 10 Types of Solids, Liquids, and Gases](#)

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[* * CHM 112 M. Prushan Why Do Solids and Liquids Exist? Intermolecular Forces gases -](#)

[Little to no attraction between molecules CHM 112 M. Prushan Why Do Solids and Liquids](#)

[Exist? Intermolecular Forces liquids - Fairly strong forces between a few molecules CHM 112](#)

[M. Prushan Why Do Solids and Liquids Exist?](#)

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CHAPTER 10: LIQUIDS AND SOLIDS - Components of a solid are close together and exert large attractive

force on each other. - Gas: low density but high compressibility - Solid: high density but slight compressibility. -

Liquid: fall somewhere between gas and solid - From the solid to liquid state, and liquid stage to gas state of water

H 2 O, there are extensive attractive forces among the molecules in liquid water, similar to but not as strong as

those in the solid state. o For water: the ...

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Chapter 10 Solids and Liquids. Opening Essay. There is an urban legend that glass is an extremely thick liquid rather than a solid, even at room temperature. Proponents claim that old windows are thicker at the bottom than at the top, suggesting that the glass flowed down over time. Unfortunately, the proponents of this idea have no credible ...

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Chapter 10 - Liquids and Solids . 10.1 Intermolecular Forces . A. Dipole-Dipole Forces 1. Attraction between molecules with dipole moments a. Maximizes (+) ----- (-) interactions b. Minimizes (+) ----- (+) and (-) ----- (-) interactions 2. About 1% of strength of ionic bonds a. Unimportant in gas phase due to distance between molecules

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The three main states of matter are solid, liquid, and gas. Plasma is the fourth state of matter.

Several exotic states also exist. A solid has a defined shape and volume. A common example is ice.

A liquid has a defined volume, but can change state. An example is liquid water. A gas has

neither a defined shape nor volume.

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State 483 10.9 Phase Diagrams 491 For Review 496 Key Terms 496 Questions and Exercises 498.

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[Chapter 10 - Liquids and Solids . III. Liquid State Phenomena. Structural models of liquids.](#)

More complex than models for solids or gases for two reasons. Liquids have strong forces. Liquids

have significant Surface. tension - the of a liquid to increase its surface area. For the surface area

of a liquid to increase, molecules would have to ...

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...