
Chapter 11 Review Gases Section 1 Answer Key

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CHAPTER 11
REVIEW Gases
SECTION 1
SHORT ANSWER
Answer the
following questions
in the space
provided. 1. b
Pressure surf f a o c
r e ce area. For a
constant force, when
the surface area is
tripled the pressure
is (a) doubled. (b) a
third as much. (c)
tripled. (d)
unchanged. 2. d, c,
a, b Rank the
following pressures
in increasing order.

(a) 50 kPa (c) 76 torr
(b) 2 atm (d) 100
N/m² 3.
SECTION 11.1
Gases and
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Chapter 11 Review
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Answers

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Section 11.4 Dalton's Law of Partial Pressures Goals To describe the properties of mixtures of gases. To describe calculations that deal with mixtures of gases. In the real world, gases are usually mixtures. This section describes how mixing gases affects the

properties of the resulting mixture.
Gas Law Problems Combined
Ideal Density, Molar Mass, Mole Fraction, Partial Pressure, Effusion Chapter 11
Liquids and Intermolecular Forces Chapter 10
Gases: Part 1 of 12
Chapter 11 - Liquids and Intermolecular Forces: Part 1 of 10
Chapter 11 Gas Laws - Day 1 - Gases
Pressure Chapter 10
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Dipole-Dipole and Hydrogen Bonding: Chapter 11 - Part 1
Chapter 11 Bankruptcy Basics Kinetic Molecular Theory and the Ideal Gas Laws
Intermolecular Forces and Boiling Points
Intermolecular forces and Boiling points
Intermolecular Forces - Hydrogen

Bonding, Dipole-Dipole, Ion-Dipole, London Dispersion Interactions
Partial Pressures
Vapor Pressure: Crash Course Chemistry #15
Chapter 10 - Gases
Chapter 10 - Gases: Part 2 of 12
Intermolecular Forces Which gas equation do I use?
11. Kinetic Theory of Gases Part 5
General Chemistry 1 Review Study Guide - IB, AP,
College Chem Final Exam
Pressure exerted by liquids and gases
chapter 11 class 8 science part 8
FSc Physics Book1,

CH 11, LEC 1: Pressure of Gases
Chapter 10 (Gases)
- Part 1 review questions chapter 11 Class 10(Physics) ~~World Climate \u0026 Climate Change~~
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Chapter 11- Gases: Section 1: Gases and Pressure ...
this theory explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample. SECTION 11.1 Key Terms
pressure newton barometer millimeters of mercury
Chapter 11 - Gases
CHAPTER 11
REVIEW Gases
SECTION 2

SHORT ANSWER
Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances:
increase a. temperature increases, volume stays the same
decrease b. volume increases, temperature stays the same
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Section 11.3 Equation Stoichiometry and Ideal Gases ...

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Combined
Ideal Density,
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Pressure, Effusion
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Chapter 11 Gas Laws

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Intermolecular forces
and Boiling points
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Dipole-Dipole, Ion-
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Interactions Partial
Pressures

Vapor Pressure: Crash
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Gases Chapter 10 -
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do I use? 11. Kinetic
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Questions

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93 Gases
CHAPTER 11
REVIEW Gases
SECTION 1
SHORT ANSWER

Answer the following questions in the space provided. 1. _____ Pressure = . For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. _____ Rank the following pressures in increasing order.

If a gas and a liquid are the same temperature and pressure, diffusion occurs

much faster in the gas because. A. there are more elastic collisions between the particles in a gas. B. gases are more compressible. C. the particles move faster in a gas and there is a greater distance between them.