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If a gas and a liquid are the same temperature and pressure, diffusion occurs much faster in the gas because. A. there are more elastic collisions between the particles in a gas. B. gases are more compressible. C. the particles move faster in a gas and there is a greater distance between them. Chemistry Section 11 Answers - Free PDF File Sharing

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Modern Chemistry 93 Gases CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. _____ Pressure = . For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. _____ Rank the following pressures in increasing order.

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Chapter 11 - Gases

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Chapter 11 Gases - An Introduction to Chemistry

462 Chapter 11 Gases Discovering the Relationships Between Properties If we want to explain why a weather balloon carrying instruments into the upper atmosphere expands as it rises, we need to consider changes in the properties of the gases (pressure, volume, temperature, or number of gas particles) inside and outside the balloon.

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Section 11.4 Dalton 's Law of Partial Pressures Goals To describe the properties of mixtures of gases. To describe calculations that deal with mixtures of gases. In the real world, gases are usually mixtures. This section describes how mixing gases affects the properties of the resulting mixture.

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this theory explains some of the properties of ideal gases. In this chapter, you will study the predictions of kineticmolecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample. SECTION 11.1 Key Te r m s pressure newton barometer millimeters of mercury

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