

## Chapter 11 Review Gases Section 3 Modern Chemistry Answers

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SECTION 1 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. b Pressure — orce For a constant force, when the surface area is tripled the surface area pressure is (a) doubled. as much. (c) tripled. 7-0 (d) unchanged. Rank the following pressures in increasing order. (c) 76 torr (a) 50 kPa O, OOİctbv-x

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part 8

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Modern Chemistry 93 Gases CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_ Pressure = . For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d)

unchanged. 2. \_\_\_\_\_ Rank the following pressures in increasing order.

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this theory explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample. SECTION 11.1 Key Terms  
pressure newton barometer millimeters of mercury

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462 Chapter 11 Gases Discovering the Relationships Between Properties If we want to explain why a weather balloon carrying instruments into the upper atmosphere expands as it rises, we need to consider changes in the properties of the gases (pressure, volume, temperature, or number of gas particles) inside and outside the balloon.

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CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER

Answer the following questions in the space provided. 1. b Pressure surf f a o c r e ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged. 2. d, c, a, b Rank the following pressures in increasing order. (a) 50 kPa (c) 76 torr (b) 2 atm (d) 100 N/m<sup>2</sup> 3.

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Section 11.4 Dalton's Law of Partial Pressures Goals To describe the properties of mixtures of gases. To describe calculations that deal with mixtures of gases. In the real world, gases are usually mixtures. This section describes how mixing gases affects the properties of the resulting mixture.

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