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Chapter 12 Stoichiometry 127

SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of

interacting moles, representative particles, masses, and gas volume at STP.

chapter 6 balancing stoich worksheet and key

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Science - Borders, Jennie / Chemistry

An expanded version of the flowchart for stoichiometric calculations illustrated in Figure 11.4.1 is shown in Figure 12.2.1. We can use the balanced

chemical equation for the reaction and either the masses of solid reactants and products or the volumes of solutions of reactants and products to determine the amounts of other species, as ...

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stoichiometry problems; one for each reagent given. If 2 products are given, pick one and use it for both calculations If 10.6 g of copper reacts with 3.83 g sulfur, how many grams of the product (copper (I) sulfide) will be formed?

Stoichiometry Section 12.1 the Arithmetic Of Equations ...

Stoichiometry Practice Worksheet Solve the following stoichiometry grams-grams problems: Using the following equation: 2 NaOH + H2S04 -+ 21-120 + Na2S04 ... Chapter 12 Stoichiometry . In the reaction represented by the equation 2Na 2H20 —y 2NaOH + H2, how many grams of

24 Stoichiometry Section 12.1 the Arithmetic Of Equations ...

Chapter 6 Balancing and Stoichiometry Worksheet and Key Topics: • Balancing Equations • Writing a chemical equation • Stoichiometry Practice: 1. In the reaction: 4Li (s) + O 2 (g) ® 2Li 2O (s) a. what is the product? b. what are the reactants? c. what does the " (s) " after the formula of lithium oxide signify?