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SECTION 12.1 THE ARITHMETIC  
OF EQUATIONS (pages 353-358)

This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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Chapter 11 Small-Scale Lab Section  
11.3 Precipitation Reactions: Formation  
of Solids, page 345 Analysis 1. a. Na  
2CO<sub>3</sub> 2AgNO<sub>3</sub> ? 2NaNO<sub>3</sub> 3Ag 2CO  
3(s) b. 2Na<sub>3</sub>PO<sub>4</sub> 3Pb(NO<sub>3</sub>)<sub>2</sub> ?  
6NaNO<sub>3</sub> Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>(s) 2.

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Stoichiometry CHAPTER 11 SOLUTIONS MANUAL Section 11.1 Defining Stoichiometry pages 368–372 Practice Problems pages 371–372 1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is

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Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question. \_\_\_\_ 1. The first step in most stoichiometry problems is to \_\_\_\_.

a. add the coefficients of the reagents  
c. convert given quantities to

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volumes b. convert given quantities to moles d.  
convert given quantities to masses \_\_\_\_ 2.

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36. a. Two formula units  $\text{KClO}_3$  decom- pose  
to form two formula units  $\text{KCl}$  and three  
molecules  $\text{O}_2$ . b. Four molecules  $\text{NH}_3$  react  
with six molecules  $\text{NO}$  to form ?ve mol-