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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows: $4Al + 3O_2 \rightarrow 2Al_2O_3$ a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3, how many moles of O_2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: $2Fe + 3H_2SO_4 \rightarrow Fe_2SO ...$

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Textbook pages: Chapter 12. Key Terms: stoichiometry. mole-mole problems. mass-mass problems. mass-volume problems. volume-volume problems. particle – particle problems. expected yield. actual yield. percent yield Directions: Use this information as a general reference tool to guide you through this unit. Don't hesitate to ask your teacher for help! By the conclusion of this unit, you ...

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Stoichiometry 379 CHAPTER 12 Assessment 36 Chapter 12 stoichiometry test b answer key. a. Two formula units KClO3 decom- pose to form two formula units KCl and three molecules O2. b Chapter 12 stoichiometry test b answer key. Four molecules NH3 react with six molecules NO to form fi ve mol-

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Quantitative calculations that involve the stoichiometry of reactions in solution use volumes of solutions of known concentration instead of masses of reactants or products. The coefficients in the balanced chemical equation tell how many moles of reactants are needed and how many moles of product can be produced.