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to form two formula  
units KCl and three  
molecules O<sub>2</sub>. b.  
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Have students  
consider the sizes of  
the containers  
shown rela-tive to  
the sizes of  
containers used in a  
classroom  
laboratory. Have  
them imag-  
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CHAPTER 12 ST  
OICHIOMETRY  
Name\_\_\_\_\_per\_\_\_\_  
12.1 What is  
Stoichiometry? 1.  
What is meant by

the use of the term  
“ quantitative ’  
in the definition of  
stoichiometry on  
page 354? 2. How  
does conservation  
of mass have to do  
with  
stoichiometry? 3.  
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12, Stoichiometry  
(continued) 8. The  
coefficients of a  
balanced chemical  
equation tell you  
the relative number

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of moles of \_\_\_\_\_  
and \_\_\_\_\_ in a  
chemical reaction. 9.  
Why is the relative  
number of moles of  
reactants and  
products the most  
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2 guided reading  
and study  
workbook. this  
section explains a  
key d. 10.  
Complete the  
table about states  
of matter. States  
of Matter. State  
Shape Volume.  
Definite 22  
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two more  
questions, in the  
table As you read,  
write answers.  
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a chemical  
reaction from  
data. •

Determine the  
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chemical reaction.  
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(pages 353 – 358)  
This section  
explains how to  
calculate the

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