

Chapter 12 Stoichiometry Guided Reading Study Work Answers

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Stoichiometry Reading Guide Study Guide

for Chapter 12 (Stoichiometry) p. 357 #2 p.
379 #61, 64, 69, 70, 73, 86, 88, 90 p. 877

Chapter 12 # 5-10 p. 880 Chapter 14 #22

Answers:

12.1 The Arithmetic of Equations 12

Chapter 12.1, 12.2 Stoichiometry p1 Unit 1

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Limiting Reactant

Stoichiometry Basic Introduction, Mole to
Mole, Grams to Grams, Mole Ratio

Practice Problems Ch 12.1- 12. 2

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Reading (Level J/ K) Introductory Tip-to-
Tail Vector Addition Problem Why is the
Sky Blue? Find the Average Atomic Mass -
Example: Magnesium

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Study Work Answers

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Chapter 3: Stoichiometry – Guided

Reading Section 3.1 – 3.2 1. True or

False? Most hydrogen atoms have a mass

of 1.008 amu. Justify your answer. If

true, explain why it is true. If false, what

mass do most hydrogen atoms have?

False, 1.008 amu is actually hydrogen ' s

average mass, NO atom of hydrogen

actually has the mass of 1.008 amu. 2.

Chapter 12 Guided Reading Stoichiometry

Answer Key

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Stoichiometry Answer Key. Chapter 12

Guided Reading Stoichiometry Chapter 12

Stoichiometry127 SECTION 12.1 THE

ARITHMETIC OF EQUATIONS (pages

353 – 358) This section explains how to

calculate the amount of reactants

required or product formed in a

nonchemical process. It teaches you how

to interpret chemical equations in terms

of interacting moles, representative

particles, masses, and gas volume at

STP.

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Chapter 12 Stoichiometry127. SECTION

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EQUATIONS

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Chapter 12 Stoichiometry127. SECTION

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Chapter 12 Stoichiometry Reading

Guide

Introduce the term sto- ichiometry in

your own words. Stress that

stoichiometry allows students to

calculate the amounts of chemical sub-

stances involved in chemical reactions

using information obtained from bal-

anced chemical equations.

Chapter 12 Stoichiometry127.

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OF EQUATIONS (pages 353 – 358)

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