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Stoichiometry Problem: Mass Precipitate

Kinetics: Initial Rates and Integrated Rate

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Stoichiometry Problem: Mass Precipitate Kinetics: Initial Rates and Integrated Rate Laws Limiting Reagent, Theoretical Yield, and Percent Yield Solution Stoichiometry

Finding Molarity, Mass \u0026amp; Volume

Stoichiometry Problem: Titration Calculation

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Chapter 12 (Chemical Kinetics) - Part 2

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1 The first step in most stoichiometry problems is Page 3/15.

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

*Chapter 12.2: Stoichiometry of Reactions in Solution ...*

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