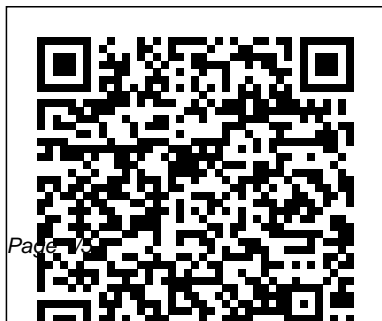


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Acceptable answers include the idea of writing a ratio using the coefficients of two substances from a balanced

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3. If 6.52 L of  $\text{H}_2$  react with excess nitrogen what volume of  $\text{NH}_3$  is produced?

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Quantitative calculations involving reactions in solution are carried out in the same manner as we discussed in Chapter 11. Instead of masses, however, we use volumes of solutions of known concentration to determine the number of moles of reactants. Whether we are dealing with volumes of solutions of reactants or masses of reactants, the coefficients in the balanced chemical equation tell us ...

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Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question. \_\_\_\_ 1. The first step in most stoichiometry problems is to \_\_\_\_.

a. add the coefficients of the reagents  
c. convert given quantities to volumes  
b. convert given quantities to moles  
d. convert given quantities to masses \_\_\_\_ 2.

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