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Answer Key

Answers 1. a. $2 \text{ }^{23}\text{Al} \text{ }^{3}\text{O} \text{ or } 2 \text{ }^{23}\text{Al} \text{ }^{2}\text{O}$
moles $\text{Al }^{2}\text{O}_3 = (2.3)(2)/4 = 1.2$ moles b. Using
the same ratios, moles $\text{O}_2 = (3.9)(3)/2 = 5.6$
moles 2. a. 2 moles Fe gives 3 moles H_2 ,
moles $\text{H}_2 = (1.7)(3)/2 = 2.6$ moles b. 3 moles
 H_2SO_4 gives 1 mole product moles yield = 3
 $\times 2.8 = 8.4$ moles 3. Mole ratios: 2 mol Mg/ 2
mol MgO = 1 mol Mg: 1 mol product 1 mol O

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Answers

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12.2.1 Construct mole ratios from balanced
chemical equations and apply these ratios in
stoichiometric calculations. 12.2.2 Calculate
stoichiometric quantities from balanced chemical
equations using units of moles, mass, representative
particles, and volumes of gases at STP. Guide for
Reading

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Answers

CHAPTER 2 REVIEW Measurements and
Calculations SECTION 2 SHORT ANSWER

Answer the following questions in the space
provided. 1. Complete the following conversions:
a. 100 mL 0.1 L b. 0.25 g 25 cg c. 400 cm³ 0.4 L d.
400 cm³ 0.0004 m 2. For each measuring device
shown below, identify the quantity measured and
tell when it would remain constant ...

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