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to which a compound dissociates in  
water. The higher the solubility  
product constant, the more soluble  
the compound. AP Chemistry  
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Electrochemistry . 18.1 Balancing Oxidation-  
Reduction Equations . A. The Half-Reaction  
method (this is a review from Unit 2, Chapter  
4) 1. Write the half-reactions 2. Balance the  
elements that are NOT oxygen and hydrogen  
3.

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Learning Objectives. Given a balanced  
equation, write the reactant quotient (Q) or  
equilibrium constant (K) Know how the  
magnitude of K is related to the extent of  
the reaction ...

### Chapter 17 - AP Chemistry

#### Chapter 17 Ap Chemistry

Chapter 4: Types of Chemical  
Reactions and Solution  
Stoichiometry 4.1 Water the Common  
Solvent 4.2 The Nature of Aqueous  
Solutions: Strong and Weak  
Electrolytes

#### Chapter 17. Additional Aspects of Equilibrium

Chapter 3; Chapter 4; Chapter  
5; Chapter 6; Chapter 7;  
Chapters 8 & 9; Chapter 10;

Chapter 11; Chapter 12;  
Chapter 13; Chapter 14;  
Chapter 15; Chapter 16;  
Chapter 17; Chapter 18;  
Chapter 22; Other; Labs  
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**Gallico, Book 6, Chapter 17 - Gallo-**  
**Roman Mythology** AP Bio Chapter 17 2  
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Hasselbalch equation  
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common-ion effect. buffer. how  
buffers work. two ways to make  
a buffer. if there is a weak  
electrolyte in equilibrium and  
you add an i... - sol'n that  
contains a weak acid-conjugate  
base or  
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the redox reac.on 2. Iden.fy: A.  
Oxidaon Half-Reac.on B. Reduc.on  
Half-Reac.on C. Oxidizing Agent D.  
Reducing Agent Zn(s)+Cu<sup>2+</sup>(aq ...  
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Chapter 17. Additional  
Aspects of Equilibrium. 17.1  
The Common Ion Effect. • The  
dissociation of a weak  
electrolyte is decreased by  
the addition of a strong  
electrolyte that has an ion.  
in common. with the weak  
electrolyte. • For example,  
consider the ionization of a  
weak acid, acetic acid. HC.  
2H.  
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Cameron*

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STUDY. PLAY. Common-ion  
effect. Whenever a weak  
electrolyte and a strong  
electrolyte contain a common  
ion, the weak electrolyte  
ionizes less than it would if  
it were alone in solution.  
Buffered solutions Acid base.

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