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Chapter 17. Additional Aspects of Equilibrium.

17.1 The Common Ion Effect. • The dissociation of a weak electrolyte is decreased by the addition of a strong electrolyte that has an ion. in common. with the weak electrolyte. • For example, consider the ionization of a weak acid, acetic acid. HC. 2H.

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### *Chapter 17. Additional Aspects of Equilibrium*

#### Chapter 4: Types of Chemical Reactions and Solution Stoichiometry

#### 4.1 Water the Common Solvent 4.2

#### The Nature of Aqueous Solutions: Strong and Weak Electrolytes

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AP Chemistry: Chapter 17. common-ion effect. buffer. how buffers work. two ways to make a buffer. if there is a weak electrolyte in equilibrium and you add an i.... - sol'n that contains a weak acid-conjugate base or

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~~- Cardiovascular Emergencies~~

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Ksp, this value indicates the

degree to which a compound

dissociates in water. The higher

the solubility product constant, the more soluble the compound. AP Chemistry Chapter 17 Review Flashcards | Quizlet I can advise you this service - ? www.HelpWriting.net ? Bought essay here. No problem.

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Learning Objectives. Given a balanced equation, write the reactant quotient (Q) or equilibrium constant (K) Know how the magnitude of K is related to the extent of the reaction ...

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Chapter 17 AP Chemistry. STUDY. PLAY. Common-ion effect. Whenever a weak electrolyte and a strong electrolyte contain a common ion, the weak electrolyte ionizes less than it would if it were alone in solution. Buffered solutions Acid base. [ap chemistry science chapter 17 Flashcards and Study Sets ...](#)

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aplikasidapodik.com Chapter 18 -

Electrochemistry . 18.1 Balancing

Oxidation-Reduction Equations . A.

The Half-Reaction method (this is

a review from Unit 2, Chapter 4)

1. Write the half-reactions 2.

Balance the elements that are NOT

oxygen and hydrogen 3.