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Chapter 17. Additional
Aspects of Equilibrium. 17.1
The Common Ion Effect. • The
dissociation of a weak
electrolyte is decreased by the
addition of a strong electrolyte
that has an ion. in common.
with the weak electrolyte. •

For example, consider the
ionization of a weak acid,
acetic acid. $\text{HC}_2\text{H}_3\text{O}_2$.
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Chapter 17. Additional Aspects of Equilibrium

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Chapter 17. Learning Objectives. Given a balanced equation, write the reactant quotient (Q) or equilibrium constant (K) Know how the magnitude of K is related to the extent of the reaction ...

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Chapter 17 AP Chemistry. STUDY. PLAY. Common-ion effect. Whenever a weak electrolyte and a strong electrolyte contain a common ion, the weak electrolyte ionizes less than it would if it were alone in solution. Buffered solutions Acid base.

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Electrochemistry Cameron

Chapter 17 Ap Chemistry Review Ksp, this value indicates the degree to which a compound dissociates in water. The higher the solubility product constant, the more soluble the compound. AP Chemistry Chapter 17 Review Flashcards | Quizlet I can advise you this service - ? www.HelpWriting.net ? Bought essay here. No problem.

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Chapter 18 –
Electrochemistry . 18.1
Balancing Oxidation-
Reduction Equations . A.
The Half-Reaction method
(this is a review from Unit
2, Chapter 4) 1. Write the
half-reactions 2. Balance
the elements that are NOT
oxygen and hydrogen 3.
**Chapter 17 Ap Chemistry
Review**

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confidence level for several
years. Such as a cultural
dynamic, womack s
approach moves falla-
cies to the mandatory inclusion
of outlines are still
mastering.

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Chapter 17—Additional

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Chemical Reactions and
Solution Stoichiometry 4.1
Water the Common Solvent
4.2 The Nature of Aqueous
Solutions: Strong and Weak
Electrolytes

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common-ion effect. buffer.
how buffers work. two ways to
make a buffer. if there is a
weak electrolyte in equilibrium
and you add an i.... - sol'n that
contains a weak acid-
conjugate base or

Chapter 17

Electrochemistry AP
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Cameron Redox Prac.ce
Consider: 1. Balance the
redox reac.on 2. Iden.fy:
A. Oxidation Half-Reac.on
B. Reduction Half-Reac.on
C. Oxidizing Agent D.
Reducing Agent
 $\text{Zn(s)} + \text{Cu}^{2+}(\text{aq}) \rightarrow \dots$