

Chapter 17 Thermochemistry Study Guide Answers

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THERMOCHEMISTRY

CHAPTER 17, Thermochemistry (continued) 2. What is calorimetry? Calorimetry is the accurate and precise measurement of heat change for chemical and physical processes. 3. Use Figure 17.5 on page 511. Circle the letter next to each sentence that is true about calorimeters.

G The calorimeter container is insulated to minimize loss of heat to or

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Name _____ Date _____ Period _____ Chapter 17 Thermochemistry Study Guide 17.1 – 17.2

Thermochemical Equations 1. Make the following conversions: a. 444 cal to joules = 1.86×10^3 J b. 1.8 kJ to joules = 1.8×10^3 J c. 0.45 kJ to calories = 1.1×10^2 cal 2. Classify each of these processes as endothermic or exothermic: a. condensing steam ...

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Workbook CHAPTER 17, Thermochemistry (continued) 6. In

thermochemical calculations, is the direction of heat flow given from the ... thermochemical equation in the first paragraph on page 517 as a guide.

SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520 – 526)

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SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505–510)

Study Guide Thermochemistry $Q = m C \Delta T$ $\Delta H = -q C = 4.184$ J moles g oC 1. When 150-g sample of KCl dissolves in 65.0 g of water in a calorimeter, the temperature

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Test. PLAY. Match. Gravity. Created by. Caleb_Tucker1. Terms in this set (55) Energy. The capacity for doing work or supplying heat. Thermochemistry. The study of energy changes that occur during chemical reactions and changes in state.

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CHAPTER ...

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Ch 17 Thermochemistry Practice Test Matching Match each item with the correct statement below.

a. calorimeter d. enthalpy b. calorie e. specific heat c. joule f. heat capacity _____ 1. quantity of heat needed to raise the temperature of 1 g of water by 1°C _____ 2. SI unit of energy _____ 3.

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CHAPTER 17, Thermochemistry (continued) GUIDED PRACTICE PROBLEM 12 (page 513) 12. When 50.0 mL of water containing 0.50 mol HCl at 22.5°C is mixed with 50.0 mL of water containing 0.50 mol NaOH at 22.5°C in a calorimeter, the temperature of the solution increased to 26.0°C.

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