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Chapter 17 Thermochemistry183 ... 184 Guided Reading and Study Workbook CHAPTER 17,Thermochemistry (continued) 6. In thermochemical calculations, is the direction of heat flow given from the ... thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520–526)

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CHAPTER 17, Thermochemistry (continued) GUIDED PRACTICE PROBLEM 12 (page 513) 12. When 50.0 mL of water containing 0.50 mol HCl at 22.5 ° C is mixed with 50.0 mL of water containing 0.50 mol NaOH at 22.5 ° C in a calorimeter, the temperature of the solution increased to 26.0 ° C.

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can you tell us about how this topic is studied and what its uses are? Take the quiz and we ' ll find out together!

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Ch 17 Thermochemistry Practice Test Matching Match each item with the correct statement below. a. calorimeter d. enthalpy b. calorie e. specific heat c. joule f. heat capacity ____ 1. quantity of heat needed to raise the temperature of 1 g of water by 1 ° C ____ 2. SI unit of energy ____ 3.

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CHAPTER 17, Thermochemistry (continued) 2. What is calorimetry? Calorimetry is the accurate and precise measurement of heat change for chemical and physical processes. 3. Use Figure 17.5on page 511. Circle the letter next to each sentence that is true about calorimeters. G The calorimeter container is insulated to minimize loss ofheat to or

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SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505 – 510)

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Name _____ Date _____ Period ____ Chapter 17 Thermochemistry Study Guide 17.1 –

17.2 Thermochemical Equations 1. Make the following conversions: a. 444 cal to joules = 1.86 x 103 J b. 1.8 kJ to joules = 1.8 x 103 J c. 0.45 kJ to calories = 1.1 x 102 cal 2. Classify each of these processes as endothermic or exothermic: a. condensing steam ...

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