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## Chapter 17 Thermochemistry Study Guide

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Thermochemistry is the  
scientific study of the heat  
energy that is involved in  
or produced by chemical  
reactions and/or physical  
transformations. What can  
you tell us about how this  
topic is studied and what its  
uses are? Take the quiz and  
we ' ll find out together!  
Chapter 17 Thermochemistry  
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CHAPTER 17,  
Thermochemistry (continued)  
GUIDED PRACTICE  
PROBLEM 12 (page 513) 12.  
When 50.0 mL of water

containing 0.50 mol HCl at  
22.5 ° C is mixed with 50.0 mL  
of water containing 0.50 mol  
NaOH at 22.5 ° C in a  
calorimeter, the temperature of  
the solution increased to  
26.0 ° C.

### **Thermochemistry & the Laws of Thermodynamics - Study.com**

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Miss\_Wenckowski includes 23  
questions covering vocabulary,  
terms and more. Quizlet  
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PLAY. Match. Gravity. Created by. Caleb\_Tucker1. Terms in this set (55) Energy. The capacity for doing work or supplying heat. Thermochemistry. The study of energy changes that occur during chemical reactions and changes in state. Chemistry Chapter 17: Thermochemistry: Study Guide ... CHAPTER 17, Thermochemistry (continued) 2. What is calorimetry? Calorimetry is the accurate and precise measurement of heat change

for chemical and physical processes. 3. Use Figure 17.5on page 511. Circle the letter next to each sentence that is true about calorimeters. G The calorimeter container is insulated to minimize loss of heat to or THERMOCHEMISTRY 5 Lessons in Chapter 17: Prentice Hall Chemistry Chapter 17: Thermochemistry Chapter Practice Test ... Study.com has thousands of articles about every imaginable degree, area of study and career ...

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15.1 – 15.2 Heat, Calorimetry,  
and Enthalpy. Make the  
following conversions: 444 cal to  
joules . ... Chapter 17  
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SECTION 17.1 THE FLOW  
OF ENERGY HEAT AND  
WORK (pages 505 – 510)  
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Test Matching Match each item  
with the correct statement below. a.  
calorimeter d. enthalpy b. calorie e.  
specific heat c. joule f. heat capacity  
\_\_\_\_ 1. quantity of heat needed to  
raise the temperature of 1 g of water  
by 1 ° C \_\_\_\_ 2. SI unit of energy  
\_\_\_\_ 3.  
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Thermochemistry  $Q = m C$   
 $T \quad H = -q C = 4.184 \text{ J}$   
moles g oC 1. When 150-g  
sample of KCl dissolves in  
65.0 g of water in a  
calorimeter, the temperature

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Prentice Hall Chemistry Chapter 17: Thermochemistry ...  
Chapter 17  
Thermochemistry 183 ... 184  
Guided Reading and Study  
Workbook CHAPTER  
17, Thermochemistry  
(continued) 6. In  
thermochemical calculations, is  
the direction of heat flow given  
from the ... thermochemical  
equation in the first paragraph  
on page 517 as a guide.  
SECTION 17.3 HEAT IN  
CHANGES OF STATE (pages  
520 – 526)  
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Name \_\_\_\_\_ Date \_\_\_\_\_ Period  
\_\_\_\_\_ Chapter 17  
Thermochemistry Study Guide  
17.1 – 17.2 Thermochemical  
Equations 1. Make the following  
conversions: a. 444 cal to joules =  
 $1.86 \times 10^3 \text{ J}$  b. 1.8 kJ to joules =  $1.8 \times 10^3 \text{ J}$  c. 0.45 kJ to calories =  $1.1 \times 10^2 \text{ cal}$   
2. Classify each of these  
processes as endothermic or  
exothermic: a. condensing steam ...  
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