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Chemistry End of Chapter Exercises. Write a brief description or definition of each of the following: (a) nucleon (b) α particle (c) β particle (d) positron (e) γ ray (f) nuclide (g) mass number (h) atomic number. Which of the various particles ( α particles, β particles, and so on) that may be produced in a nuclear reaction are actually ...

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The production of energy in a nuclear reactor can be stopped by pulling out all control rods. A breeder reactor produces more fuel than it uses. The fission products produced in nuclear power plants are not radioactive. An uncontrolled chain reaction led to the nuclear accident in Chernobyl, Ukraine. Chemistry: Matter and Change Chapter 25 149

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Chapter 19 Radioactivity and Nuclear Energy 1. The nucleus of an atom has little or ... 18. The elements with atomic number greater than 92 are referred to as the transuranium elements. The transuranium elements have been prepared by bombardment reactions of other nuclei. 19. 13 27Al + 2 4 He 15 30 P + 0 1n 20. The probe of a Geiger (Geiger-M ü ller) counter contains contains argon gas. The ...

*Chapter 19 Radioactivity and Nuclear Energy*

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**Chapter 18 NuClear Chemistry**

\ Holt Chemistry NY: Chapter 18 - Nuclear Chemistry. Holt Chemistry NY: Chapter 18 – Nuclear Chemistry Flashcard. nucleons. a proton or a neutron. The subatomic particles that are in the nucleus. nuclide. an atom that is identified by the number of protons and neutrons in its nucleus. ex: " Carbon-14 " strong force. the interaction that binds nucleons together in a nucleus. It is the force ...

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**Chapter 18 Nuclear Chemistry Answer**

718 Chapter 18 Nuclear Chemistry + + + + Energy p + p + n + n 24 He2+ For many of the lighter elements, the possession of an equal number of protons and neutrons leads to stable atoms. For example, carbon-12 atoms, 6C 12, with six protons and six neutrons, and oxygen-16 atoms, 8O 16, with eight protons and eight neutrons, are both very stable ...

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Nuclear Chemistry Nuclear Transformations • Rutherford in 1919 performed the first nuclear transformation. • The transmutations are sometimes represented by listing in order, the target nucleus, the bombarding particle, the ejecting particle and the product nucleus. • The above equation becomes: 14 2 17 1 7 4 8 1N + He O + H→ 14 17

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Figure 18.19 A " Fossil Nuclear Reactor " in a Uranium Mine Near Oklo in Gabon, West AfricaMore than a billion years ago, a number of uranium-rich deposits in West Africa apparently " went critical, " initiating uncontrolled nuclear fission reactions that may have continued intermittently for more than 100,000 years, until the concentration of uranium-235 became too low to support a chain ...

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Chapter 18 Nuclear Chemistry. Chapter Map. Nuclides • Nuclide = a particular type of nucleus, characterized by a specific atomic number and nucleon number • Nucleon number or mass number = the number of nucleons (protons and neutrons) in the nucleus of a nuclide. Nuclide Symbolism. Nuclear Stability • Electrostatic force = the force that causes opposite electrical charges to attract each ...

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Chapter 18 – Nuclear Chemistry 289 Key Ideas Answers 14. Because protons and neutrons reside in the nucleus of atoms, they are called nucleons. 16. There are two forces among the particles within the nucleus. The first, called the electrostatic force, is the force between electrically charged particles. The second force,