

Chapter 21 Quantization Of Energy

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Chapter 21 Quantization Of Energy

Energy is quantized in some systems, meaning that the system can have only certain energies and not a continuum of energies, unlike the classical case. This would be like having only certain speeds at which a car can travel because its kinetic energy can have only certain values. We also find that some forms of energy transfer take place with discrete lumps of energy.

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Apply the Wilson-Sommerfeld quantization rules, and show that the possible values of the total energy are predicted to be $E = n^2 h^2 / 2I$ where You would like to observe an E. coli bacterium that ...

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So we know that physics got turned upside down at the turn of the 20th century, but how did that all begin? What was the first thing that set this whole chain...

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Quantization. It may be helpful at this point to further consider the idea of quantum states.

Atoms, molecules, and fundamental electron and proton charges are all examples of phy

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Max Planck postulated that energy was quantized and could be emitted or absorbed only in integral multiples of a small unit of energy, known as a quantum. The energy of a quantum is proportional to

the frequency of the radiation; the proportionality constant h is a fundamental constant (Planck's

Chapter 21 Quantization Of Energy

Quantization of energy also resembles a standing wave on a string that allows only particular harmonics described by integers. It is also similar to going up and down a hill using discrete stair steps rather than being able to move up and down a continuous slope. Your potential energy takes on discrete values as you move from step to step.

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