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Thermodynamics An Engineering Approach

The change in internal energy can be found from the first law of thermodynamics:  $U = Q - W = (3.5 \times 10^3 \text{ J}) - (2.1 \times 10^3 \text{ J}) = 0.9 \times 10^3 \text{ J} = 90 \text{ kJ}$ . A gas in a cylinder is kept at a constant pressure of  $3.5 \times 10^5 \text{ Pa}$  while 300 kJ of heat are added to it, causing the gas to expand from  $0.9 \text{ m}^3$  to  $1.5 \text{ m}^3$ .

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C. 1/4 of the volume is in liquid point and the remaining at vapor form. Determine the

pressure of the steam, and quality of the saturated mixture, and density of the

mixture. Given: Volume (V) Temperature (T) C. Find: The pressure of the steam.