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MANUAL 9. Evaluate the
experiments that led to the
conclu-sion that electrons
are negatively charged
particles found in all
matter. The deflection toward
positively charged plates

demonstrated the negatively
charged nature of electrons;
the fact that changing the
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The solubility rules referenced in
the following answers are from
Table 41 of the text The phrase
“slightly soluble” is interpreted to
mean insoluble and the phrase
“marginally soluble” is interpreted
to mean soluble a Soluble (Rule 3)
b

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Sr(NO₃)₂ SrSO₄ Sr(HSO₄)₂

Sr(H₂PO₄)₂ SrO SrCl₂ NH₄ NO₃ (NH₄)₂ SO₄ NH₄ HSO₄ NH₄ H₂ PO₄

(NH₄)₂ O NH₄ Cl Al(NO₃)₃ Al₂(SO₄)₃ Al(HSO₄)₃ Al(H₂ PO₄)₃

Al₂ O₃ AlCl₃

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Column A
1. The rays and particles that are emitted by a radioactive material
2. A reaction that involves a change in an atom's nucleus
3.

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4 Cl(s). How many grams of NH₃ (g) reacted? Is the law of conservation of mass observed in the above reaction? Justify your answer. 157.5 g 106.5 g 51 g Yes. Mass of reactants equals mass of products. Section 3.2 Assessment page 79 10. Classify each example as a physical change or a chemical change. a.

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Chapter Assessment Answer Key Chemistry: Matter and Change T167 Name Date Class 20 Understanding Main Ideas Chemistry: Matter and Change • Chapter 4 Chapter Assessment (Part A) Use the periodic table to identify each element described

below. 1. atomic number 65 2. 78 protons 3. 44 protons and 44 electrons 4. atomic number 24 5. 21 protons 6. [9th Class Chemistry, Ch 4, Exercise Ch no 4-Matric Part 1 Chemistry Class 9 MCQ Question Answer of Science Chapter-4 Structure of the Atom for CBSE Exam |](#)

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