

Chapter 6 Review Chemical Bonding Answer Key

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CHAPTER 6 Chemical Bonding

Modern Chemistry 41 Chemical Bonding CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. _____ A chemical bond between atoms results from the attraction between the valence electrons and _____ of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2.

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Youâ€™™ need to be familiar with three types of chemical bonds for the SAT II Chemistry exam: ionic bonds, covalent bonds, and metallic bonds. Chapter 6: Chemical Bonding - GO.HRW.COM go.hrw.com/hrw.nd/gohrw_rls1/pKeywordResults?HC2%20CH06 Chapter 6: Chemical Bonding: Welcome to Modern Chemistry. Chapter 6 describes covalent bonds, ionic bonds, Lewis dot structures, resonance structures, and â€!
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...
an attraction between atoms that allows the formation of chemical substances that contain two or more atoms. The bond is caused by the electrostatic force of attraction between opposite charges, either between electrons and nuclei, or as the result of a dipole attraction.

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Section 6.2 – Covalent Bonding A covalent bond is a chemical bond in which two atoms share a pair of valence electrons.

When two atoms share one pair of electrons, the bond is called a

single bond. Covalent vs Ionic Bond

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ANSWER Answer the following questions in the space provided. 1. a
A chemical bond between atoms results from the attraction between
the valence electrons and of different atoms. (a) nuclei (c) isotopes (b)
inner electrons (d) Lewis structures 2. b A covalent bond consists of
(a) a shared electron.

6 Chemical Bonding

Chemical bonding that results from the electrical attraction
between positive ions and negative ions is called ionic bonding.
In a purely ionic bond, the metal atom gives up its electron or
electrons to the nonmetal atom. In covalent bonding, a bond
forms from the sharing of electron pairs between two atoms. In a
purely covalent bond, the

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CHAPTER 6 REVIEW Chemical Bonding SECTION 5 SHORT
ANSWER Answer the following questions in the space provided. 1.
Identify the major assumption of the VSEPR theory, which is used to
predict the shape of atoms. Pairs of valence electrons repel one
another. 2. In water, two hydrogen atoms are bonded to one oxygen
atom. Why isn ' t water a linear molecule?

Chapter 6 – Chemical Bonds

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electrical attraction between positive ions and negative ions is
called ionic.

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between atoms results from the attraction between the valence
electrons and _____ of different atoms. nuclei. A covalent bond
consists of.