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2 6.1 (a) The Clausius-Clapeyron equation: $dP_{sat} / P_{sat} = \frac{h_{i,vap}}{RT^2} dT$ or $\ln P_{sat} = -\frac{h_{i,vap}}{R} \left(\frac{1}{T} - \frac{1}{373} \right) + C$ so $P_{sat} = 101 \dots$

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Having said that in Class 11, students will come across the topic in Chapter 6 of the chemistry textbook and they have to examine a lot of concepts related to thermodynamics. They will be learning about the interchange between heat and work, laws of thermodynamics, calorimetry, reaction enthalpy, etc.

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