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in this set (20) Monatomic Ion. Consists of a single atom with a
positive or negative charge resulting from the loss of gain of one
or more valence

Chapter 9 – Chemical Names and Formulas

Writing the Names and Formulas for Bases. A base is a
compound that produces OH⁻ in water. When naming a base,
you name it like any other compound that starts with a regular
or transition metal. Ex: NaOH = sodium hydroxide. When
writing the formula for a base, remember to balance charges.
Ex: magnesium hydroxide = Mg(OH)₂

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Formulas Answer Section MATCHING 1. ANS: A PTS: 1 DIF:
L1 REF: p. 253 OBJ: 9.1.1 Identify the charges of monatomic
ions by using the periodic table, and name the ions. STA: Ch.5.a
2.ANS: H PTS: 1 DIF: L1 REF: p. 254

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How to Write Chemical Formulas from Compound Names

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and Molecular Formula Introduction Covalent Compounds:

Writing Chemical Names and Formulas

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Chemical Names and Formulas 281 CHAPTER 9 Assessment 42. a.
2- b. 1+ c. 1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+ 44. a. barium ion b.
iodide ion c. silver ion d. mercury(II) ion 45. cyanide, CN⁻ and
hydroxide, OH⁻ 46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d.
oxide ion 47. zero; A compound is electrically neutral. 48. The
symbols for the cation and

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First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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Chemical Names and Formulas

To name a binary molecular compound, use the following guidelines:

1. Write the names of the elements in the order listed in the formula 2.

Use prefixes appropriately to indicate the number of each kind of

atom 3. End the name of the second element with the suffix -ide

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SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC

COMPOUNDS I. Write the formulas for these binary ionic compounds. c.

potassium iodide 2. write the formulas for the compounds formed c. b. f. 3.

Name the following binary ionic compounds. sodium sulfide g. CuCl: h.

sncl. l. a. MnO₂ c. d. SrBr₂ 221 Chapter 9 Chemical Names and Formulas

Chapter 9 Chemical Names And Formulas Workbook Answers

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Monatomic Ion. Consists of a single atom with a positive or

negative charge resulting from the loss of gain of one or more

valence electrons and behaves as a unit, yet it still has a charge.

(+ or -). Determined by the amount of electrons lost or gained.

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1. when the name of the anion (X) ends in -ide the acid name begins with

the prefix hydro-. The stem of the anion has the suffix -ic and is followed

by the word acid (ex. HCl= hydrochloric acid) 2. when the anion name

ends in -ite the acid name is the stem of the anion with the suffix -ous

followed by the word acid (ex. H₂SO₃ -- X=sulfite = sulfurous acid)

Name Date Class CHEMICAL NAMES AND FORMULAS

9

Chapter 9 "Chemical Names and Formulas" Tools. Copy this to

my account; E-mail to a friend; Find other activities; Start over;

Help; Use these activities to help yourself to learn the

vocabulary and major topics presented in this chapter. A B;

anion: any atom or group of atoms that has a negative charge:

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