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Stoichiometry

CHAPTER 9 REVIEW

Stoichiometry

SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. _____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂

Chapter 9 Review
Stoichiometry Mixed
Chemistry Worksheet on
Stoichiometry Mixed
Review. Assume all reactions go to completion. Write the formula equation, balance the equations, and solve the problems. Draw a rectangle around the answer and don't forget the units. Methane (CH₄) combines with oxygen to form carbon dioxide and water. Balanced equation:

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**Chemistry Worksheet on
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Stoichiometry CHAPTER 9
REVIEW Stoichiometry MIXED
REVIEW SHORT ANSWER
Answer the following questions
in the space provided. 1. Given
the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ _____ a. What is the value
of the coefficient x in this
equation? _____ b. What is the
molar mass of C₃H₄?

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Stoichiometry b.

Theoretically, how many moles of NH_3 will be produced? **PROBLEMS** Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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Modern Chemistry 2

Stoichiometry CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following equation: C ...

Modern Chemistry 3

Stoichiometry MIXED REVIEW continued c. If 0.1 mol of N 2 combine with H 2, what must be true about the quantity of H

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Stoichiometry SECTION 3

PROBLEMS Write the

answer on the line to the

left. Show all your work in the space provided. 1. 88%

The actual yield of a

reaction is 22 g and the theoretical yield is 25 g.

Calculate the percentage

yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$...

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REVIEW. SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + x O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What ...

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ANSWER Answer the

following questions in the space provided. 1. Modern chemistry chapter 9 3 review stoichiometry answers. Download:

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Stoichiometry SECTION

9-3 PROBLEMS Write the

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N_2 are mixed with 12.0 mol

of H_2 according to the

following equation: $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

$3\text{H}_2(\text{g}) + 2\text{NH}_3(\text{g}) \rightarrow$

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((Stoichiometry)) Flashcards |
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Stoichiometry 2. CH 9 Limiting
Reactant . CH 9 Pract. Limit.

Yield . CH 9 Percent Yield-New.

CH 9 Percent-Actual-Theoretical

Yield. CH 9 Section Review 9.1

- 9.3 . CH 9 Pretest . CH 9 Test

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other study tools. ... What is the

limiting reactant in this reaction

when 0.750 mol of N_2H_4 is

mixed with 0.500 mol of

H_2O_2 ? .750 mol $\text{N}_2\text{H}_4 \times 2$ mol

$\text{H}_2\text{O}_2 = 1.5$ mol H_2O_2 ... Chapter

8: Chemistry 18 Terms. idkkerin.

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Stoichiometry CHAPTER 9

REVIEW Stoichiometry

SECTION 3 PROBLEMS

Write the answer on the line to

the left. Show all your work in

the space provided. 1. _____

The actual yield of a reaction

is 22 g and the theoretical

yield is 25 g. Calculate the

percentage yield. 2. 6.0 mol of

N_2 are mixed with 12.0 mol

of H

Chapter 9 Stoichiometry

Test Answer Key Modern

Chemistry

mixed with 9.0g H_2 . What is

the theoretical yield in grams

of the product? $\text{N}_2 + 3\text{H}_2 \rightarrow$

2NH_3 ; 17 Application. ...

Chapter 9 Review

Stoichiometry - Chapter 9

Review. Stoichiometry.

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9 - Review ... Limiting reagent.
Theoretical yield.
Stoichiometry. Percent yield.