

## Chemactivity 4 Answer Key

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4. The IE of Li<sup>+</sup> should be larger than the IE of He because both atoms have 2 electrons in the 1st shell and Li<sup>+</sup> has a core charge of +3 whereas He only has a core charge of +2. 5. The IE of F<sup>-</sup> should be less than the IE of Ne because both atoms have eight electrons in the 2nd shell and F<sup>-</sup> has a core charge of +7 whereas Ne has a core charge ...  
7.11: The Activity Series - Chemistry LibreTexts  
PLEASE NOTE: If you have a question about these answers, it is your responsibility to come to office hours or ask during class work time. UNIT 12 - HW Practice Keys - ChemActivity 51: Cell Voltage - ChemActivity 50: Electrochemical Cell UNIT 11 - HW Practice Keys - ChemQuest 55: Free Energy - ChemQuest 54: 2nd Law of Thermodynamics  
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3. What is the mass (in grams) of 4.35 ! 106 atoms of <sup>12</sup>C? 4. What is the mass (in grams) of 6.022 ! 1023 atoms of <sup>12</sup>C? 5. What is the mass (in grams) of one molecule of methane which has one <sup>12</sup>C atom and four <sup>1</sup>H atoms, <sup>12</sup>C<sup>1</sup>H<sup>4</sup>? 6. a) Define mass number. b) Define atomic number. 4  
ChemActivity 1 The Nuclear Atom  
[Chem 30A Additional Resources - Cabrillo College](#)  
Chemactivity 4 Answer Key (See chemactivity 5 for explanation) Two atoms have the same nuclear charge, but one atom has a lower IE than the

other atom. How could this happen? The first atom must have a greater distance between the valence electron and the nucleus than the other atom. \*\*different distances change results.

[Chemactivity 26 Answers - Exam Answers Free](#)

Chemactivity 4 Answer Key (See chemactivity 5 for explanation) Two atoms have the same nuclear charge, but one atom has a lower IE than the other atom. How could this happen? The first atom must have a greater distance between the valence electron and the nucleus than the other atom.

\*\*different distances change results.

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PhaseChanges\_IMF\_activity\_answers.pdf - 86 kB; Download all files as a compressed .zip. Title ChemActivity: Phase Changes and Intermolecular Forces: Description Students use the States of Matter simulation to identify phase changes at the particulate level. They compare water and oxygen to identify why the phase changes occur at different ...

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(See chemactivity 5 for explanation) Two atoms have the same nuclear charge, but one atom has a lower IE than the other atom. How could this happen? The first atom must have a greater distance between the valence electron and the nucleus than the other atom. \*\*different distances change results.

**Answers to Exercises and Problems**

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*Chem; GI; Answers*

ChemActivity 30 Limiting Reagent Information 177 Chemists refer to the reactant which limits the amount of product that can be made from a given collection of original reagents as the limiting reagent or limiting reactant. Critical Thinking Questions 3. Identify the limiting reagent for CTQ 2. 4.

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**Chemistry: A Guided Inquiry Critical Thinking Questions ...**  
Activity Answer Keys: These are the keys to the Activities found in the  
laboratory Manual as well as some others that were given out in class.  
These keys are to be used for you to self-assess after completion of the  
assignments.

### *The Nuclear Atom*

Nuclear Chemistry Part A and Part B HW Questions Assigned  
as HW on 12/4/18 and 12/5/18. Nuclear Chemistry Part A

Questions Answer Key. Nuclear Chemistry Part A and Part B  
HW Questions Answer Key. \*\*There is a mistake in the Part B  
answer Key. #9 should read; 32 means: The mass number.

### **Chemactivity 4: The Shell Model (I) Flashcards | Quizlet**

4. 5. 6. How many trunks are found in one dozen elephants? Give your  
answer in terms of a number (such as 17 or  $3.25 \times 10^{15}$  trunks). How  
many legs are found in one dozen elephants? Give your answer in terms of  
a number (such as 17 or  $3.25 \times 10^{15}$  legs). How many carbon atoms are  
found in one dozen methane, CH<sub>4</sub>, molecules? Give

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ChemActivity 25 Metals 145 4. Explain, in terms of  
electronegativities, why the dividing line between metals and  
nonmetals in the periodic table (the red line that steps down  
from beneath B to between Po and At) is oriented the way that it  
is (as opposed to being, for example, horizontal or

vertical). Model 2: The Electronic Structure of Metals ...

[Chemistry a Guided Inquiry Pages 151 - 200 - Flip PDF ...](#)

ChemActivity 4 1. 1.31 MJ/mole 2. The electron that is farthest from the  
nucleus will have the least negative potential energy (d is larger) and the  
lowest ionization energy. 3.

### **HW Keys - Roosevelt High School AP Chemistry 2017-18**

4. The answer here depends on the students' answer to CTQ 3.

5.a) The IE 1 for He is greater than the IE 1 for H because the  
nuclear charge on a He atom is +2 whereas the nuclear charge  
on H is +1. b) The IE 1 for Li is less than the IE 1 for He  
because at least one of the electrons of Li must be farther away  
from the nucleus than any electron ...

[Chem:GI:Answers - Weebly](#)

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