

# Chemfiesta Balancing Equations Simple Stoichiometry Answers

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## Stoichiometry Homework Answer Key Chemfiesta

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The first step to balance the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation. There is an arrow between the sides, signaling the direction the reaction is happening in.

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Stoichiometry Review Worksheet. 1)Using the  
following equation:  $2 \text{NaOH} + \text{H}_2 \text{SO}_4 \rightarrow 2 \text{H}_2 \text{O} + \text{Na}_2 \text{SO}_4$  How many grams of sodium sulfate  
will be formed if you start with 200 grams of  
sodium hydroxide and you have an excess of  
sulfuric acid ( $\text{H}_2 \text{SO}_4$ )? 2)Using the following  
equation:  $\text{Pb}(\text{SO}_4)_2$

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Stoichiometry Review Worksheet. 1)Using the  
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Stoichiometry-KEY Balance the following  
equations: 1)  $1 \text{N}_2 + 3 \text{F}_2 \rightarrow 2 \text{NF}_3$  2)  $2 \text{C}_6 \text{H}_{10} + 17 \text{O}_2 \rightarrow 12 \text{CO}_2 + 10 \text{H}_2 \text{O}$  3)  $1 \text{HBr} + 1 \text{KHCO}_3 \rightarrow 1 \text{H}_2 \text{O} + 1 \text{KBr} + 1 \text{CO}_2$  4)  $2 \text{GaBr}_3 + 3 \text{Na}_2 \text{SO}_3 \rightarrow 1 \text{Ga}_2 (\text{SO}_3)_3 + 6 \text{NaBr}$  5)  $3 \text{SnO} + 2 \text{NF}_3 \rightarrow 3 \text{SnF}_2 + 1 \text{N}_2 \text{O}_3$  Using the  
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KEY Balance the following equations: 1)  $1 \text{N}_2 + 3 \text{F}_2 \rightarrow 2 \text{NF}_3$  2)  $2 \text{C}_6 \text{H}_{10} + 17 \text{O}_2 \rightarrow 12 \text{CO}_2 + 10 \text{H}_2 \text{O}$  3)  $1 \text{HBr} + 1 \text{KHCO}_3 \rightarrow 1 \text{H}_2 \text{O} + 1 \text{KBr} + 1 \text{CO}_2$  4)  $2 \text{GaBr}_3 + 3 \text{Na}_2 \text{SO}_3 \rightarrow \dots$  and the reagent  
that leads to this answer is the limiting reagent.

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