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The first step to balance the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation.

There is an arrow between the sides, signaling the direction the reaction is happening in. [Click here to access this Book](#)

Balancing Equations and Simple

Stoichiometry-KEY

Balance the following equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$ 2) 2 C_6

$\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ 3) $1 \text{ HBr} + 1 \text{ K}_2\text{CO}_3 \rightarrow 1 \text{ KBr} + 1 \text{ CO}_2$ 4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$ 5) $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3$

Using the following equation: $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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following equation: $2 \text{NaOH} +$
 $\text{H}_2 \text{SO}_4 \rightarrow 2 \text{H}_2 \text{O} + \text{Na}_2$
 SO_4 How
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Review Worksheet. 1)Using the
following equation: $2 \text{NaOH} +$
 $\text{H}_2 \text{SO}_4 \rightarrow 2 \text{H}_2 \text{O} + \text{Na}_2 \text{SO}_4$. 4
How many grams of sodium
sulfate will be formed if you start
with 200 grams of sodium
hydroxide and you have an excess
of sulfuric acid ($\text{H}_2 \text{SO}_4$)?
2)Using the following equation:
 $\text{Pb}(\text{SO}_4)$.
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Stoichiometry Review
Worksheet. 1)Using the
following equation: $2 \text{NaOH} +$
 $\text{H}_2 \text{SO}_4 \rightarrow 2 \text{H}_2 \text{O} + \text{Na}_2$

SO. 4 How many grams of sodium sulfate will be formed if you start with

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Stoichiometry-KEY Balance the

following equations: 1) $1 \text{ N}_2 + 3$

$\text{F}_2 \rightarrow 2 \text{ NF}_3$ 2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2$

$\rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ 3) 1 HBr

$+ 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr}$

$+ 1 \text{ CO}_2$ 4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2$

SO_3 ... and the reagent that

leads to this answer is the limiting

reagent. Both Balancing

Equations and Simple

Stoichiometry-KEY

Balancing Equations and

Simple Stoichiometry-KEY

Balance the following

equations: 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2$

NF_3 2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow$

$12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ 3) 1 HBr

$+ 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1$

$\text{KBr} + 1 \text{ CO}_2$ 4) $2 \text{ GaBr}_3 + 3$

$\text{Na}_2\text{SO}_3 + 1 \text{ Ga}_2(\text{SO}_3)_3 + 6$

$\text{NaBr} \rightarrow 5 \text{ SnO} + 2 \text{ NF}_3 + 3 \text{ SnF}_2$

$+ 1 \text{ N}_2\text{O}_3$ Using the

following equation: $2 \text{ NaOH} +$

$\text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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