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Balancing Equations and Simple Stoichiometry-KEY
 Balance the following equations:
 1) $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
 2) $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
 3) $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$
 4) $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$
 5) $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1$

N 2 O 3 Using the following
equation: $2 \text{NaOH} + \text{H}_2 \text{SO}_4$
 $2 \text{H}_2 \text{O} + \text{Na}_2 \text{SO}_4$

Stoichiometry

Worksheet Chemfiesta

Answers

Balancing Equations
and Simple

Stoichiometry-KEY

Balance the following
equations: 1) $1 \text{N}_2 +$

$3 \text{F}_2 \rightarrow 2 \text{NF}_3$ 2) 2C_6

$\text{H}_{10} + 17 \text{O}_2 \rightarrow 12 \text{CO}_2$

$+ 10 \text{H}_2 \text{O}$ 3) $1 \text{HBr} +$

$1 \text{KHCO}_3 \rightarrow 1 \text{H}_2 \text{O} + 1$

$\text{KBr} + 1 \text{CO}_2$ 4) 2GaBr

$3 + 3 \text{Na}_2 \text{SO}_3 \rightarrow 1 \text{Ga}_2$

$(\text{SO}_3)_3 + 6 \text{NaBr}$ 5) 3

$\text{SnO} + 2 \text{NF}_3 \rightarrow 3 \text{SnF}_2 +$

$1 \text{N}_2 \text{O}_3$ Using the

following

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Worksheet. 1)Using the

following equation: $2 \text{NaOH} +$

$\text{H}_2 \text{SO}_4 \rightarrow 2 \text{H}_2 \text{O} + \text{Na}_2 \text{SO}_4$

4 How many grams of sodium

sulfate will be formed if you

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Worksheet. 1)Using the

following equation: $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
SO. 4 How
Chemfiesta Stoichiometry
Practice Answers
The first step to balance the
equation is to write down the
chemical formula of reactants
that are listed on the left side of
the chemical equation. After
this, you can list down the
products on the right hand side
of the chemical equation.
There is an arrow between the
sides, signaling the direction
the reaction is happening in.
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Balancing Equations Answer

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 Worksheet. 1) Using the
 following equation: $2 \text{NaOH} +$
 $\text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
 4 How many grams of sodium
 sulfate will be formed if you
 start with 200 grams of sodium
 hydroxide and you have an
 excess of sulfuric acid (H_2SO_4)?
 2) Using the following
 equation: $\text{Pb}(\text{SO}_4)_2$
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 Practice Answers
 Balancing Equations and Simple
 Stoichiometry-KEY Balance the
 following equations: 1) $1 \text{N}_2 + 3$
 $\text{F}_2 \rightarrow 2 \text{NF}_3$ 2) $2 \text{C}_6\text{H}_{10} + 17 \text{O}_2$
 $\rightarrow 12 \text{CO}_2 + 10 \text{H}_2\text{O}$ 3) 1HBr

+ $1 \text{KHCO}_3 + 1 \text{H}_2\text{O} + 1 \text{KBr} +$
 1CO_2 4) $2 \text{GaBr}_3 + 3 \text{Na}_2\text{SO}_3$
 $\rightarrow 3 \text{Ga}_2(\text{SO}_3)_3 + 6 \text{NaBr}$ 5) 3
 $\text{SnO} + 2 \text{NF}_3 \rightarrow 3 \text{SnF}_2 + 1 \text{N}_2\text{O}$
 3 Using the following equation: 2
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} +$
 Na_2SO_4
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 Practice Worksheet Answers
Balancing Equations and
Simple Stoichiometry-KEY
 Balancing Equations and
 Simple Stoichiometry-KEY
 Balance the following
 equations: 1) $1 \text{N}_2 + 3 \text{F}_2 \rightarrow$
 2NF_3 2) $2 \text{C}_6\text{H}_{10} + 17 \text{O}_2$
 $\rightarrow 12 \text{CO}_2 + 10 \text{H}_2\text{O}$ 3) 1
 $\text{HBr} + 1 \text{KHCO}_3 \rightarrow 1 \text{H}_2\text{O}$

+ $1 \text{KBr} + 1 \text{CO}_2$ 4) 2GaBr_3
 $+ 3 \text{Na}_2\text{SO}_3 \rightarrow \dots$ and the
 reagent that leads to this
 answer is the limiting
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