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Chapter 6 Balancing and Stoichiometry Worksheet and Key

Topics: • Balancing Equations • Writing a chemical equation • Stoichiometry Practice: 1. In the reaction:  $4\text{Li (s)} + \text{O}_2 \text{ (g)} \rightarrow 2\text{Li}_2\text{O (s)}$  a. what is the product? b. what are the reactants? c. what does the “ (s) ” after the formula of lithium oxide signify?

**Balancing Equations and Simple**

**Stoichiometry-KEY**

And this is where the art of balancing chemical equations starts to come into play. The general idea is, Try to balance the... try to balance the molecules that have multiple elements in them first, and leave the... molecules that only have one element in them for last.

[Balancing Equations | The](#)

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Balancing equations: Learn how to fill out all of those little blanks on the endless papers your teacher has been giving you.

Writing complete equations:

It ' s not enough to balance equations – you also have to know how to add more information to make their interpretation easier. Find out how to add the symbols of state and little Greek letters to your equations.

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earlier. This is the mole ratio  
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## Tessshebaylo

A chemistry tutorial designed to help learn the basic principles of balancing chemical equations, along with examples and methods of balancing different chemical equations.

### Stoichiometry and Balancing Reactions - Chemistry LibreTexts

Balancing Equations and Simple Stoichiometry-KEY  
Balance the following equations: 1)  $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$  2)  $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$  3)  $1 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + 1 \text{ CO}_2$  4)  $2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2(\text{SO}_3)_3 + 6 \text{ NaBr}$  5)  $3 \text{ SnO} + 2 \text{ NF}_3 \rightarrow 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3$  Using the following equation:  $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

*chapter 6 balancing stoich worksheet and key*

The way that we balance equations is to write coefficients in front of the chemical formulas in an equation. These coefficients multiply the number of atoms of each element in these formulas, so if you see a "2" in front of water, you'll know you have four hydrogen atoms and two oxygen atoms.

### Simple Balance Equations Worksheets - Lesson Worksheets

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