

# Chemfile Mini Guide To Problem Solving

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 1. Carry out the following computations, and express the result in sci-enti?c notation: a.  $2.290 \times 10^7 \text{ cm} \times 4.33 \times 10^3 \text{ s}$  ans:  $5.29 \times 10^3 \text{ cm/s}$  b.  $1.788 \times 10^5 \text{ L} \times 7.111 \times 10^3 \text{ m}^2$  ans:  $2.514 \times 10^3 \text{ L/m}^2$  c.  $5.515 \times 10^4 \text{ L} \times 6.04 \times 10^3 \text{ km}$  ans:  $9.13 \text{ L/km}$  d.  $3.29 \times 10^4 \text{ km} \times 1.48 \times 10^2 \text{ min}$  ans:  $2.22 \times 10^2 \text{ km/min}$

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING The measurement has two significant figures. The measurement has four significant figures. Decimal point is present, so these zeros are significant. The rules are summarized in the following flowchart: qq Nonzero digits 3 4 0 0 g of sulfur pp Nonzero digits 4 0 0 0. mL of oxygen pp

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... Name Date Class 1 of 13. CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 8. Stoichiometry. So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of moles and that they re- late moles of a substance to mass by using the molar mass. In addition, you have learned to write chemical equations so that they represent the rearrangements of atoms that take place during chemical reactions, and you have learned to balance these equations.

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Mini Book Tracy Evans

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Mini Book Tracy Evans

I discovered the Mini-Guide to Problem Solving during my third year of teaching high school chemistry. I have found it invaluable for providing both practice problems and test problems. Each section includes sample problems with step-by-step solutions and a large selection of practice problems at various difficulty levels with the answers given in the back of the book.

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CHEMFILE MINI-GUIDE TO PROBLEM SOLVING General Plan for Solving Molarity Problems Mass of solute in g 1 Amount of solute in mol M moles solute liter solution 2 Volume of solution in L 3 Molar concentration, M 4 Convert using the molar mass of the solute. MOLARITY Molarity is the most common way to express concentration in chemistry. [CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 2...](#) CHEMFILE MINI-GUIDE TO PROBLEM SOLVING Neither pressure nor temperature is constant; therefore, the combined gas law must be used to calculate the volume of CO<sub>2</sub> at the final temperature and pressure. multiply by the volume ratio, CO<sub>2</sub> use the combined gas law to convert from the initial temperature and pressure to the final temperature and pressure C 2 H 4

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... Name Date Class 4 of 13 CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 3. Aluminum will react with sulfuric acid in the following reaction.  $2\text{Al}(s) + 3\text{H}_2\text{SO}_4(l) \rightarrow \text{Al}_2(\text{SO}_4)_3(aq) + 3\text{H}_2(g)$  a. How many moles of  $\text{H}_2\text{SO}_4$  will react ans: 27 mol  $\text{H}_2\text{SO}_4$  with 18 mol Al? b. How many moles of each product ans: 27 mol H will be produced? 2 9 mol  $\text{Al}_2(\text{SO}_4)_3$  4.

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING 1. Determine the percentage composition of each of the following compounds: a. sodium oxalate,  $\text{Na}_2\text{C}_2\text{O}_4$  ans: 34.31% Na, 17.93% C, 47.76% O b. ethanol,  $\text{C}_2\text{H}_5\text{OH}$  ans: 52.13% C, 13.15% H, 34.72% O c. aluminum oxide,  $\text{Al}_2\text{O}_3$  ans: 52.92% Al, 47.08% O d. potassium sulfate,  $\text{K}_2\text{SO}_4$  ans: 44.87% K, 18.40% S, 36.72% O 2.