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$K_{\text{Wact}} = 125 \text{ mol}$ The solution is $T_f = 549.39 \text{ K}$ c_p $\text{J mol}^{-1} \text{K}^{-1}$ solution The actual is work 499.14 K . is 25% Then greater h (b) Repeat the calculation with a temperature-dependent heat capacity $10^{-2} T - 3499$ $10^{-5} T^2 + 7.464$ $10^{-9} T^3$. . $CP(T) = 22.243 + 5977$ Assuming reversibility $T_f = 479.44 \text{ K}$.

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$2 \text{ 3 energy J N m kg m power} = = = = \text{time s s s charge current} = \text{time}$

$\text{charge} = \text{current} \cdot \text{time} = \text{A s energy power} = = \text{current} \cdot \text{electric potential}$

$\text{time } 2 \text{ 3 energy kg m electrical potential} = = \text{current} \cdot \text{time A s electrical}$

$\text{potential current} = \text{resistance } 2 \text{ 23}$

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