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At 30 degC the vapor pressure of water is 4.241 kPa. For evaluation of .1.Hp we number species as above. Llip := R·MCPH· (1773.15 - 303.15) From Example 4.7: The vapor pressure of water at 50 degC (exit of heat exchanger) is 12.34 kl'a, and water must condense to lower its partial pressure to this value.

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Chapter 1 - Section A - Mathcad Solutions. 1.4 The equation that relates deg F to deg C is: $t(F) = 1.8 t(C) + 32$. Solve this equation by setting $t(F) = t(C)$. Guess solution: $t = 0$ Given $t = 1.8t + 32 =$ Find $t() = 40$ Ans. 1.5 By definition: $P F A = F \text{ mass } g =$ Note: Pressures are in gauge pressure. $P = 3000 \text{ bar}$ $D = 4 \text{ mm}$ $A = S = 4$
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Step 1 of 4 Assume 500 ml of water in the glass and 500 ml of wine in the glass. Also assume 1 teaspoon takes 50 ml of liquid each time. Step 2 of 4 • Transfer 50 ml of water to the glass of wine and mix thoroughly.
• Now, the contaminated wine has some amount of water particle.

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