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## Chemical Quantities Chapter 10 Quiz

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*Chemical Quantities -  
AP Biology  
the SI unit  
representing  $6.02 \times 10^{23}$  representative  
particles of a*

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substance: Avogadro's  
number:  $6.02 \times 10^{23}$   
particles: standard  
temperature and  
pressure (00 C, 1 atm)  
the temperature and  
pressure at which one  
mole of gas occupies a  
volume of 22.4 L: molar  
volume: volume of a gas  
that contains one mole  
of the gas, is 22.4 L  
at STP: Avogadro ...

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CHAPTER 10: Chemical  
Quantities. CHAPTER 10:  
Chemical Quantities  
BASICS: • The basic unit

that is used to determine the  
amount of a chemical  
substance is called a mole •  
A mole(mol) of a substance  
is equivalent to  $6.02 \times 10^{23}$   
particles of that substance •  
The mole was founded by a  
scientist named Avagadro,  
and he decided to use the  
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terms and concepts of  
chemical quantities  
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Chemistry  
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Chemical Quantities  
BASICS: • The basic unit  
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substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

**CHAPTER 10  
CHEMICAL  
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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to

$6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

Prentice Hall Chemistry Chapter 10: Chemical Quantities ...

Use the chemical formula to find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles in one mole. atom  $6.02 \times 10^{23}$  O  $2 \times 6.02 \times 10^{23}$  ion Na<sup>+</sup>  $6.02 \times 10^{23}$  formula unit NaCl  $6.02 \times 10^{23}$   $6.02 \times 10^{23}$  representative particles

of a substance molecule formula unit atom

**Chemistry Chapter 10  
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*Chapter 10 Chemical Quantities Quiz Answer*

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## **Chemical Reactions & Quantities Chapter Exam - Study.com**

Chemical reactions relate quantities of reactants and products. Chemists use the mole unit to represent  $6.022 \times 10^{23}$  things, whether the things are atoms of elements or molecules of compounds. This number, called Avogadro's number, is important because this number of atoms or molecules has the same mass in grams as one atom or