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MATTER (pages 287–296) This*

section defines the mole and explains how the mole is used to measure matter.

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10: Chemical Quantities

BASICS: • The basic

unit that is used to

determine the amount of

a chemical substance is

called a mole • A

mole(mol) of a substance

is equivalent to 6.02×10^{23}

particles of that

substance • The mole

was founded by a

scientist named Avagadro,

and he decided to use the

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Chemical Quantities Practice

Problems Answer Key Use the

chemical formula to find the

number of atoms in one

molecule and multiply this

number by Avogadro ' s

number, the number of

particles in one mole. atom

6.02×10^{23} O $2 \times 6.02 \times 10^{23}$

ion Na⁺ 6.02×10^{23}

formula unit NaCl 6.02×10^{23}

6.02×10^{23}

representative particles of a

substance

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chemical formula to
find the number of
atoms in one
molecule and
multiply this
number by
Avogadro's number,
the number of
particles in one
mole. atom 6.02 ?
1023 0 2 6.02 ? 10

23 ion

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taking the measure of work a guide to validated scales ... Chemical Quantities Test Answer Key Download Chapter 10 Chemical Quantities Study Guide Answer Key - Jan 09, 2014 • 92 Guided Reading and Study Workbook CHAPTER 10, Chemical Quantities (continued) 9 Complete the table about representative particles and moles The Mass of a Mole of an Element (pages 293-294) 10 What is the atomic mass of an element? 11 Circle the letter of the phrase that completes this sentence correctly Chemical Quantities Chapter 10 Test Answer Key Download Free Chapter 10 Chemical Quantities Answer Key Chapter 10 Chemical Quantities Answer Key This is likewise one of the factors by obtaining the soft documents of this chapter 10 chemical quantities answer key by online. You might not require more epoch to spend to go to the book foundation as skillfully as search for them. *Chapter 10 Chemical Quantities Practice Problems Answer Key* Chapter 10 Chemical Quantities Test Review Answer Key

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Answers

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